**General Final Exam Review – KMT and Gases**

1) If I have 77 moles of a gas at a temperature of 25° and a pressure of 8.0 atm, what is the volume of this gas? R = 0.08206 Latm/molK.

2) If I were to compress a sample of oxygen gas from 5 L to 2 L, what would the pressure of the gas be if its initial pressure was 1.0 atm?

3) If I were to perform the process in #2 with nitrogen gas rather than oxygen gas, what would the pressure of the gas be after squishing it?

4) Why do we assume that ideal gas molecules don’t have any volume?

5) Is the concept of an ideal gas a good one or a bad one, based on what you know of the? Explain.